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Uniform B-C-N Ternary Monolayer from Non-Metal Filled g-C₃N₄ SheetNa Zhang^a, Xiao-jun Wu^{a,b,c,*}

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By using first principles calculations, four two-dimensional B-C-N ternary sheets with ordered and uniform element distribution are predicted based on the C, B, or N filled g-C₃N₄ sheet. These B-C-N ternary sheets are metallic except for B₄-C₃N₄ monolayer, which is a semiconductor with an energy band gap of 1.18 eV. In particular, the B₃C-C₃N₄ is a ferromagnetic metal with a net magnetic moment of 0.57 μ_B/cell, which can be used to develop metal-free spintronic device. The calculated formation energy indicates these B-C-N ternary sheets are highly thermal stable. It presents a new route to obtain uniform B-C-N ternary sheet for electronic and spintronic applications.

Key words: Density functional theory, B-C-N ternary monolayer, g-C₃N₄

I. INTRODUCTION

Two-dimensional (2D) materials with single atomic thickness have attracted extensive attention for their unique electronic, optical and planar properties with respect to the bulk materials [1, 2]. In particular, two analogs, *i.e.* graphene characterized as a semi-metal or zero gap semiconductor with unique electronic properties [3–5] and hexagonal boron nitride sheet (*h*-BN) as an insulator with a bandgap larger than 5 eV [6, 7], are expected to play a key role in future nanotechnology as well as to provide potential applications in next generation electronics.

In recent years, the hybrid 2D materials by combining graphene and *h*-BN monolayer, *i.e.* 2D B-C-N ternary sheets, have induced rising interest for their novel electronic and magnetic properties, remarkably different from two pristine hosts [8–16]. For instance, introducing *h*-BN flakes into graphene's framework opens a band gap of graphene, induces magnetism or half-metallicity, makes them suitable for electronic and spintronic applications [11]. Experimentally, there have been some successes in preparing 2D B-C-N compositions [8–10]. Ci *et al.* synthesized large-area atomic layers of B-C-N nanosheets, consisting of hybridized, randomly distributed domains of *h*-BN and C phases with compositions ranging from pure BN to pure graphene [8]. Raidongia *et al.* have obtained B-C-N nanosheets by

the reaction of high-surface-area activated charcoal with a mixture of boric acid and urea at 900 °C [9], and Qin *et al.* synthesized B_{0.38}C_{0.27}N_{0.35} compounds with few atomic layers by microwave plasma CVD (MPCVD) [10]. Theoretically, a lot of other B-N-C ternary systems such as BC₂N [12], BC₄N [13], B₃CN₃ [14], B₃C₂N₃ [15], have been proposed. However, it is found that most of these ternary systems have strong preferences for much more B–N bonds and C–C bonds compositions and the B-C-N ternary sheet prefers a hybridization of graphene and *h*-BN domains [8]. How to obtain B-C-N ternary monolayer material with uniform or ordered distribution of B, C, and N is still a challenge.

Graphitic carbon nitride (g-C₃N₄) is a stacked two-dimensional sheets of tris-triazine connected via tertiary amines, which is an analogue to graphite with nitrogen substitution. Due to its novel optical and chemical properties, g-C₃N₄ has been demonstrated as a potential candidate for optoelectronic conversion and photocatalytic water splitting [16–23]. Recently, ultrathin g-C₃N₄ sheets with enhanced photoresponsivity and possible magnetism have been reported experimentally, implying the great potential applications of 2D g-C₃N₄ sheet [19, 20]. Compared with graphene in structure, the g-C₃N₄ monolayer is porous sheet with uniformly distributed vacancy at atomic level [24, 25]. Thus, it provides a great potential to filling the pores in g-C₃N₄ sheet with B, C, or N atoms to prepare 2D B-C-N ternary materials with uniform and ordered distribution of B, C, and N elements.

In this work, we predicted four 2D B-C-N ternary materials by filling g-C₃N₄ sheet with C, B, and N atoms. On the basis of first principles calculations, the designed

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B-C-N sheets were either metallic or semiconductor. In particular, the B₃C-C₃N₄ sheet was ferromagnetic metal, which can be used for developing metal-free spintronic devices.

II. METHODS

All calculations are performed by using the density functional theory (DFT) method within the Perdew-Burke-Ernzerhof (PBE) generalized gradient approximation (GGA) implemented in Vienna *ab initio* Simulation Package (VASP) [26, 27]. The projector augmented wave (PAW) potential and the plane-wave cut-off energy of 500 eV are used [28, 29]. The periodic boundary condition was applied and we used a vacuum space of 15 Å along the *z* direction to avoid interactions between two B_{*x*}C_{*y*}N_{*z*}-C₃N₄ images in the nearest-neighbor unit cells. When performing the geometric optimization, both the cell shapes and atomic positions are fully relaxed for all considered B_{*x*}C_{*y*}N_{*z*}-C₃N₄ frameworks. The Brillouin zone is represented through Monkhorst-Pack special *k*-point meshes [30] with 5×5×1 and 13×13×1 for the geometry optimization and calculating the density of states. More than 20 K points along each high-symmetry line in the Brillouin zone were used to obtain the accurate band structure. The convergence criteria for total energy and force are set to be 500 eV, 1×10⁻⁴ eV, and -0.01 eV/Å, respectively. As PBE method always underestimates the band gap, the screened hybrid HSE06 functional method is used to evaluate the band gap of semiconductor systems.

III. RESULTS AND DISCUSSION

We first studied the geometric and electronic properties of the pristine 2D g-C₃N₄ framework. Our g-C₃N₄ model is based on the experimentally well characterized structure as shown in Fig.1(a) [31]. The optimized lattice constant of g-C₃N₄ unitcell is calculated to be 7.14 Å, agreeing with previous studies [24, 25]. The energy band structures with PBE and HSE06 methods indicate that the pristine g-C₃N₄ sheet is nonmagnetic semiconductor. The PBE method presents an underestimated band gap of about 1.25 eV, and this value is 2.71 eV with HSE06 method (Fig.1(b)), which is consistent well with the experimental value (2.7 eV) [31]. The orbital projected DOS analysis shows that the valence band edge is mainly contributed by N2p orbitals, whereas the conducting band edge is mainly contributed by C2p and N2p orbitals.

Next, boron, carbon, or nitrogen atoms are incorporated into the porous pristine g-C₃N₄ sheet to get uniform 2D B-C-N ternary monolayer. Four models with the compositions of B₃C-C₃N₄, B₃N-C₃N₄, B₃-C₃N₄, and B₄-C₃N₄, are designed, as shown in Fig.2. A struc-

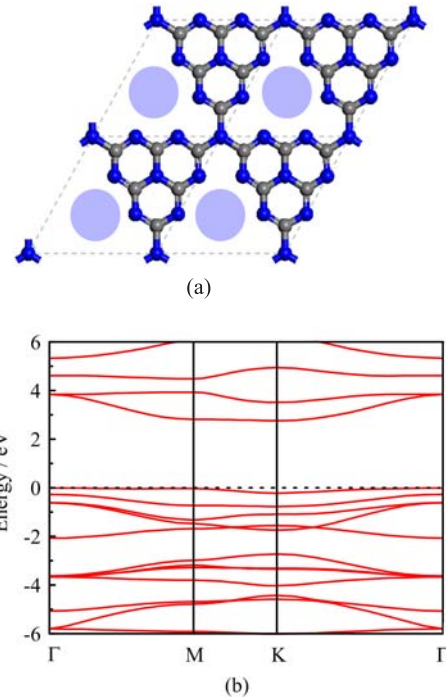


FIG. 1 (a) Geometric structure and (b) band structure of the 2D g-C₃N₄ framework calculated by HSE06. Gray and blue balls represent C and N atoms, respectively.

TABLE I Optimized lattice constant *a*, binding energy *E*_{form}, total magnetic moment per unit cell *M*, and band gap *E*_{gap} of the uniform B-C-N frameworks.

	<i>a</i> /Å	<i>E</i> _{form} /eV	<i>M</i> /μ _B	<i>E</i> _{gap} /eV
g-C ₃ N ₄	7.14	0.294	0.00	2.17
B ₃ C-C ₃ N ₄	7.27	-0.027	0.57	0.00
B ₃ N-C ₃ N ₄	7.24	-0.073	0.00	0.00
B ₃ -C ₃ N ₄	7.24	0.222	0.00	0.00
B ₄ C-C ₃ N ₄	7.21	0.080	0.00	1.18

ture distortion can be found in B₃C-C₃N₄, B₃N-C₃N₄, and B₄-C₃N₄, whereas B₃-C₃N₄ monolayer has an exactly planar structure. The optimized lattice constants are summarized in Table I. These values are slightly larger than that of pristine g-C₃N₄ monolayer.

To investigate the structure stability of proposed 2D B-C-N ternary monolayer, the average formation energies of pristine g-C₃N₄ and four B-C-N ternary monolayer models are calculated. The formation is defined as

$$E_{\text{form}} = \frac{E_{\text{tot}} - n_{\text{B}}\mu_{\text{B}} - n_{\text{C}}\mu_{\text{C}} - n_{\text{N}}\mu_{\text{N}}}{n_{\text{B}} + n_{\text{C}} + n_{\text{N}}} \quad (1)$$

where *E*_{tot} are the total energies of pristine g-C₃N₄ or B-C-N ternary monolayer. The *n*_B, *n*_C, *n*_N are the total number of B, C, and N atoms. The μ_B (-5.97 eV), μ_C (-9.22 eV), and μ_N (-8.32 eV) are the theoretical chemical potentials for B, C, and N atoms, which are

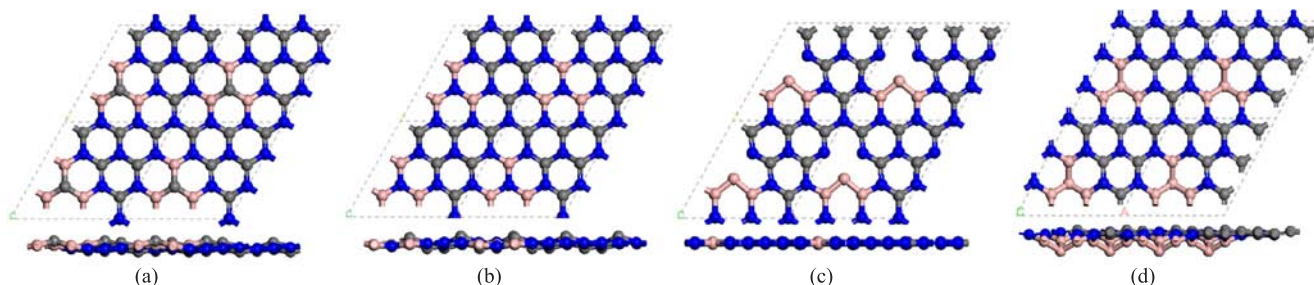


FIG. 2 The top (upper) and front (lower) views of (a) $B_3C-C_3N_4$, (b) $B_3N-C_3N_4$, (c) $B_3-C_3N_4$, and (d) $B_4-C_3N_4$. Gray, beige, and blue balls represent C, B, and N atoms, respectively.

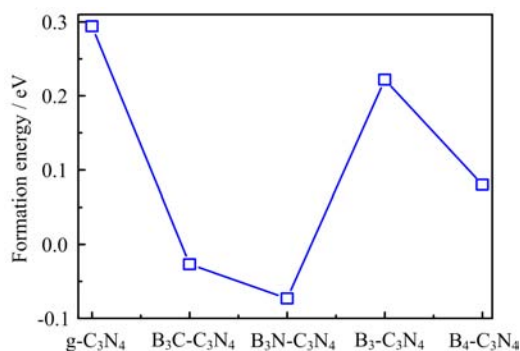


FIG. 3 The calculated average formation energy of pristine $g-C_3N_4$ and four B-C-N ternary monolayers.

obtained from diamond boron phase, graphene and $\alpha-N_2$ phase of gaseous nitrogen, respectively. The calculated average formation energy of pristine $g-C_3N_4$ and B-C-N ternary monolayer are summarized in Fig.3. It can be concluded that inserting B, C, or N atoms into the pores of $g-C_3N_4$ reduces its formation energy and these 2D B-C-N ternary monolayers are thermally stable. The $B_3N-C_3N_4$ sheet presents the highest stability in the considering systems, since the B–N bond has much higher bonding energy than B–C bond, which is consistent with the previous reports [8, 13]. The $B_3-C_3N_4$ and $B_4-C_3N_4$ systems show relatively low stability due to small bonding energy of B–B bond. The high thermal stability of $B_3N-C_3N_4$ and $B_3C-C_3N_4$ implies that they are the most probable structures to be formed in a synthesis process.

The electronic band structures of B-C-N ternary sheets are calculated within PBE and HSE06 functional schemes. The calculations indicate that $B_3N-C_3N_4$, $B_3-C_3N_4$, and $B_4-C_3N_4$ sheets are nonmagnetic. The calculated energy band gaps and magnetic moment are summarized in Table I. Figure 4 displays the electronic band structures of $B_3N-C_3N_4$, $B_3-C_3N_4$, and $B_4-C_3N_4$ sheets. Different from pristine $g-C_3N_4$, the $B_3N-C_3N_4$ and $B_3-C_3N_4$ monolayers are metallic. The $B_4-C_3N_4$ semiconducts with an energy band gap of 1.18 eV, which is much smaller than that of pristine $g-C_3N_4$. The reduced band gaps of $B_4-C_3N_4$ indicate its advantages in

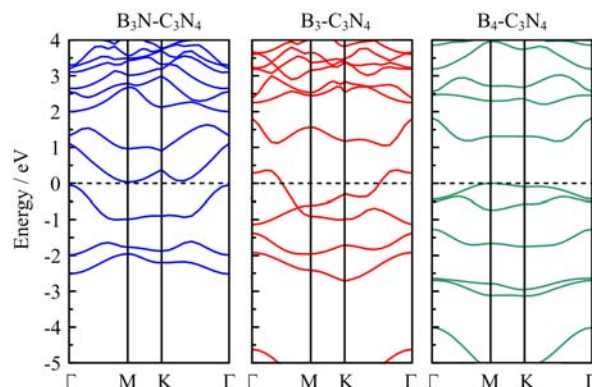


FIG. 4 The band structures of $B_3N-C_3N_4$, $B_3-C_3N_4$, and $B_4-C_3N_4$ are calculated with HSE06 method. The Fermi energy is set to be zero.

the photocatalytic reaction.

However, the $B_3C-C_3N_4$ system is magnetic. In order to explore the magnetic ground state, ferromagnetic (FM), antiferromagnetic (AFM), and nonmagnetic (NM), configurations have been considered using a (2×2) supercell. It is found that the FM state is the most energetically stable, with energy of 36.47 and 37.29 meV lower than that of the AFM and NM configurations, respectively. The calculated magnetic moment is $0.57 \mu_B/\text{cell}$ for a FM state. The calculated electronic band structure and spin charge density distribution profiles of $B_3C-C_3N_4$ system is displayed in Fig.5.

From Fig.5(a), it can be concluded that the magnetic moment of $B_3C-C_3N_4$ is mainly contributed by the carbon atoms in the system. When three boron atoms and one carbon atom are added into the porous of the pristine $g-C_3N_4$, three neighboring carbon atoms are spin polarized and the system becomes the metal-free magnetic material. Figure 5(a) also shows that the magnetic moment mainly comes from the $2p_z$ orbitals of the three neighboring carbon atoms and other atoms are slightly spin polarized. The calculated electronic band structures and atom-projected DOS (PDOS) (as shown in Fig.6(d)) analysis indicate there is strong spin-splitting for $2p_z$ orbitals of C atoms. Only one band crosses the

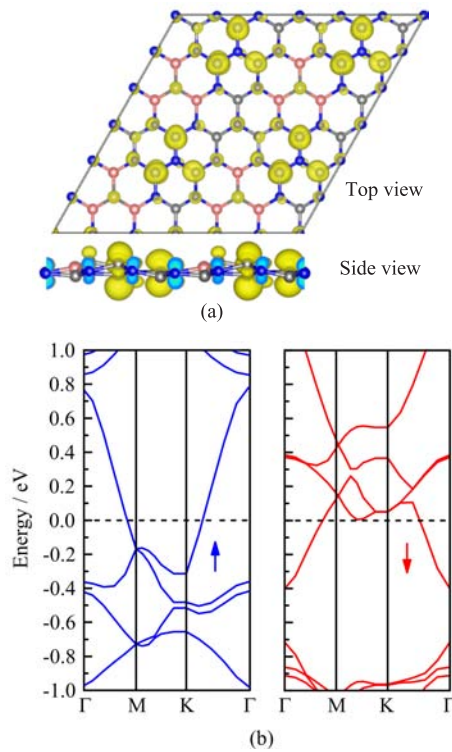


FIG. 5 (a) The top and side view of spin charge density of B₃C-C₃N₄. The isosurface is 0.002. (b) The electronic band structures are plotted for spin up and down channel with HSE06 method, respectively. The Fermi energy level is set to be zero.

Fermi level monolithically in spin up and down channels, corresponding to the C2p_z orbitals. As mentioned above, for the pristine g-C₃N₄, the valence band edge is mainly contributed by N2p orbitals, whereas the conducting band edge is mainly contributed by C2p and N2p orbitals. However, for the B₃C-C₃N₄, as can be seen, the N2p states overlap significantly with those of C2p states near the Fermi level, suggesting a strong interaction between them, which results in the spin splitting of C2p_z orbitals. The magnetism characteristics of B₃C-C₃N₄ monolayer indicate its potential applications in spintronics devices.

For the nonmagnetic systems, Fig.6 displays the calculated PDOS with PBE method. The analysis of the PDOS demonstrates that, near the Fermi level, the most significant contribution for the charge transport in such systems are associated with the p orbitals of carbon and nitride while boron atoms make a little contribution due to only one electron in p orbital.

IV. CONCLUSION

In summary, we have proposed four B-C-N ternary sheets with uniform element distribution. It is demonstrated that all of the predicted new uniform B-C-N ternary systems are metals except the B₄-C₃N₄ which

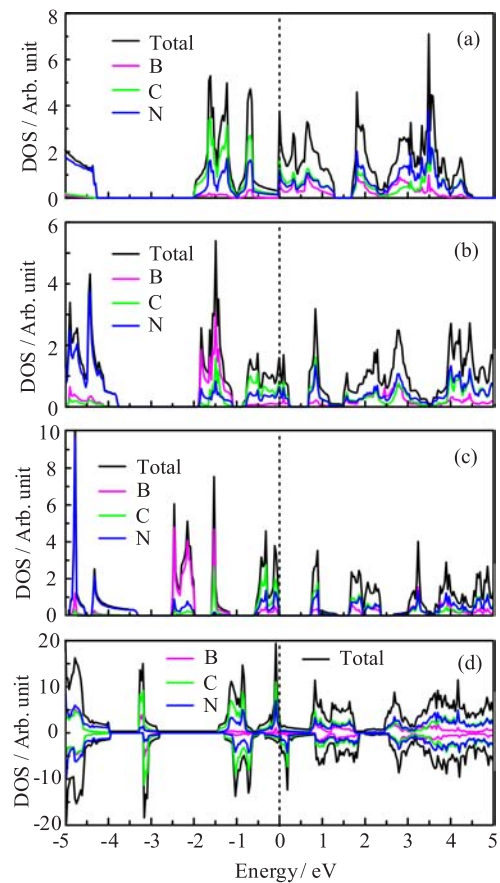


FIG. 6 The calculated DOS of (a) B₃N-C₃N₄, (b) B₃-C₃N₄, (c) B₄-C₃N₄, and (d) B₃C-C₃N₄. The Fermi energy level is set to be zero.

is a semiconductor with a band gap of 1.18 eV. Interestingly, B₃C-C₃N₄ is a metal-free ferromagnetic material and holds a magnetic moment of 0.57 μ_B /cell. The calculated formation energy indicates these B-C-N ternary sheets are highly thermal stable. Our studies present a method to design B-C-N ternary monolayer with uniform element distribution for electronic and spintronic applications.

V. ACKNOWLEDGMENTS

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- [1] C. Lee, Q. Li, W. Kalb, X. Z. Liu, H. Berger, R. W. Carpick, and J. Hone, *Science* **328**, 76 (2010).
- [2] K. S. Novoselov, D. Jiang, F. Schedin, T. J. Booth, V. V. Khotkevich, S. V. Morozov, and A. K. Geim, *Proc. Natl. Acad. Sci. USA* **102**, 10451 (2005).
- [3] J. Zhou, Q. Wang, Q. Sun, and P. Jena, *Phys. Rev. B* **81**, 085442 (2010).
- [4] E. Almahmoud, I. Kornev, and L. Bellaiche, *Phys. Rev. B* **81**, 064105 (2010).
- [5] A. Nag, K. Raidongia, K. P. S. S. Hembram, R. Datta, U. V. Waghmare, and C. N. R. Rao, *ACS Nano* **4**, 1539 (2010).
- [6] L. Song, L. Ci, H. Lu, P. B. Sorokin, C. Jin, J. Ni, A. G. Kvashnin, D. G. Kvashnin, J. Lou, B. I. Yakobson, and P. M. Ajayan, *Nano Lett.* **10**, 3209 (2010).
- [7] A. V. Krasheninnikov, P. O. Lehtinen, A. S. Foster, P. Pyykkö, and R. M. Nieminen, *Phys. Rev. L* **102**, 126807 (2009).
- [8] L. J. Ci, L. Song, C. H. Jin, D. Jariwala, D. X. Wu, Y. J. Li, A. Srivastava, Z. F. Wang, K. Storr, L. Balicas, F. Liu, and P. M. Ajayan, *Nature Mater.* **9**, 430 (2010).
- [9] K. Raidongia, A. Nag, K. P. S. S. Hembram, V. U. Waghmare, R. Datta, and C. N. R. Rao, *Chem. Eur. J.* **16**, 149 (2010).
- [10] L. Qin, J. Yu, S. Y. Kuang, C. Xiao, and X. D. Bai, *Nanoscale* **4**, 120 (2012).
- [11] A. K. Manna and S. K. Pati, *J. Phys. Chem. C* **115**, 10842 (2011).
- [12] S. Azevedo, *Phys. Lett. A* **351**, 109 (2006).
- [13] A. Freitas, S. Azevedo, M. Machado, and J. R. Kaschny, *Appl. Phys. A* **10**, 1007 (2012).
- [14] J. Y. Li, D. Q. Gao, X. N. Niu, M. S. Si, and D. S. Xue, *Nanoscale Res. Lett.* **7**, 624 (2012).
- [15] M. S. C. Mazzoni, R. W. Nunes, S. Azevedo, and H. Chacham, *Phys. Rev. B* **73**, 073108 (2006).
- [16] Y. J. Zhang, T. Mori, L. Niu, and J. H. Ye, *Energy Environ. Sci.* **4**, 4517 (2011).
- [17] Y. J. Zhang, T. Mori, J. H. Ye, and M. Antonietti, *J. Am. Chem. Soc.* **132**, 6294 (2010).
- [18] Q. J. Xiang, J. G. Yu, and M. Jaroniec, *J. Phys. Chem. C* **115**, 7355 (2011).
- [19] X. D. Zhang, X. Xie, H. Wang, J. J. Zhang, B. C. Pan, and Y. Xie, *J. Am. Chem. Soc.* **135**, 18 (2013).
- [20] D. Q. Gao, Q. Xu, J. Zhang, Z. J. Yang, M. S. Si, Z. L. Yang, and D. S. Xue, *Nanoscale* **6**, 2577 (2014).
- [21] X. F. Chen, J. S. Zhang, X. Z. Fu, M. Antonietti, and X. C. Wang, *J. Am. Chem. Soc.* **131**, 11658 (2009).
- [22] Z. X. Ding, X. F. Chen, M. Antonietti, and X. C. Wang, *ChemSusChem* **4**, 274 (2011).
- [23] X. C. Wang, X. F. Chen, A. Thomas, X. Z. Fu, and M. Antonietti, *Adv. Mater.* **21**, 1609 (2009).
- [24] M. Deifallah, P. F. McMillan, and F. Cora, *J. Phys. Chem. C* **112**, 5447 (2008).
- [25] A. Du, S. Sanvito, Z. Li, D. W. Wang, Y. Jiao, T. Liao, Q. Sun, Y. H. Ng, Z. H. Zhu, R. Amal, and S. C. Smith, *J. Am. Chem. Soc.* **134**, 4393 (2012).
- [26] J. P. Perdew, K. Burke, and M. Ernzerhof, *Phys. Rev. Lett.* **77**, 3865 (1996).
- [27] G. Kresse and J. Furthmüller, *Phys. Rev. B* **54**, 11169 (1996).
- [28] P. E. Blochl, *Phys. Rev. B* **50**, 17953 (1994).
- [29] G. Kresse and D. Joubert, *Phys. Rev. B* **59**, 1758 (1999).
- [30] H. J. Monkhorst and J. D. Pack, *Phys. Rev. B* **13**, 5188 (1976).
- [31] X. C. Wang, K. Maeda, A. Thomas, K. Takanebe, G. Xin, J. M. Carisson, K. Domen, and M. Antonietti, *Nat. Mater.* **8**, 76 (2009).